Empirical and Molecular Formulas CTR

Instructions

* Class pairs up with a person sitting next to them.
* Each pair will be given either an empirical formula card or a molecular formula card.
* Pairs with the molecular formula card will stay sitting at their tables. Pairs with the empirical formula card will stand up and search around the tables for their matching molecular formula pair. When they correctly find their matching pair (empirical + molecular), the two pairs will work together in a group of four to answer the questions on the back of the molecular formula card.

Question 1

1. Determine the molecular formula of a compound with empirical formula CH2 and molar mass of a compound 84.18 g/mol.
2. Determine the percentage composition of C and H in empirical formula.

Questions 2

1. Determine the molecular formula of a compound with empirical formula CH2O and molar mass of a compound 60.05 g/mol.
2. Determine the percentage composition of C, H, and O in empirical formula.

Question 3

1. Determine the molecular formula of a compound with empirical formulaC2H4O and molar mass of a compound 132.16 g/mol.
2. Determine the percentage composition of C, H, and O in empirical formula.

Question 4

1. Determine the molecular formula of a compound with empirical formulaC3H8 and molar mass of a compound 88.19 g/mol.
2. Determine the percentage composition of C and H in empirical formula.

Question 5

1. Determine the molecular formula of a compound with empirical formula CHand molar mass of a compound 78.11 g/mol.
2. Determine the percentage composition of C and H in empirical formula.

Question 6

1. Determine the molecular formula of a compound with empirical formulaC3H4O3 and molar mass of a compound 176.12 g/mol.
2. Determine the percentage composition of C, H, and O in empirical formula.

Question 7

1. Determine the molecular formula of a compound with empirical formulaC3H5O3 and molar mass of a compound 445.35 g/mol.
2. Determine the percentage composition of C, H, and O in empirical formula.